



Personal Protective Equipment

Understanding Solvents and PPE for Chemical Safety

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Solvents are indispensable across various industries, from chemical processing and pharmaceuticals to aviation maintenance and manufacturing. These substances are critical in cleaning spray guns, degreasing parts, and preparing surfaces for coating. Solvents range from everyday substances like water to more hazardous chemicals such as acetone, benzene, and toluene, which pose significant health risks if not handled properly. The **OSHA PPE guidelines** and **Code of Federal Regulations on Hazardous Waste** highlight how widespread solvent use affects millions of workers, underscoring the need for stringent safety measures.

Common solvents like ethanol, acetone, and ethyl acetate are used extensively in industries ranging from cleaning and pharmaceuticals to windshield de-icing and paint thinning. While these chemicals offer specialized benefits, they can pose serious health and environmental risks without adequate protective measures. MCR Safety offers an extensive range of chemical-resistant gloves and PPE solutions to protect workers from solvent exposure. Products like their **CP14R** butyl rubber gloves and **5319** Nitrile-Chem™ nitrile gloves provide superior resistance to aggressive solvents, ensuring worker safety without compromising productivity.

This overview highlights the indispensable role of solvents in modern industries and emphasizes the critical importance of proper PPE selection. By combining regulatory compliance with industry-leading protective gear, businesses can create safer work environments to protect their most valuable asset—their people.

What is a solvent?

Solvents are substances that dissolve other materials, known as solutes, to create solutions. They play a vital role in numerous industries, from cleaning and manufacturing to chemical processing and pharmaceuticals. Whether cleaning spray guns, removing grease, or preparing surfaces, solvents help keep operations running smoothly. They range from everyday substances like water to hazardous chemicals like methanol and ethyl acetate.

Because **solvents** are so widespread in industrial and consumer products, millions of workers are exposed to potentially harmful chemicals yearly. Without proper handling and protection, these solvents can cause skin irritation, respiratory issues, and long-term health problems. That's why it is critical to understand how to handle them safely and choose the right personal protective equipment

(PPE).

Types of Solvents: Polar vs. Nonpolar

A good rule of thumb is “like dissolves like”—polar solvents dissolve polar substances, while nonpolar solvents dissolve nonpolar substances. Solvents are generally classified into two types based on how they interact with different materials:

- **Polar Solvents:** These dissolve water-based or ionic substances like salts and sugars. Common polar solvents include water, ethanol, and acetone. They're widely used in cleaning, pharmaceuticals, and laboratory settings.
- **Nonpolar Solvents:** These are best for dissolving oils, grease, and hydrocarbons. Solvents like benzene, toluene, and hexane fall into this category and are commonly used in paint thinners, adhesives, and industrial degreasers.

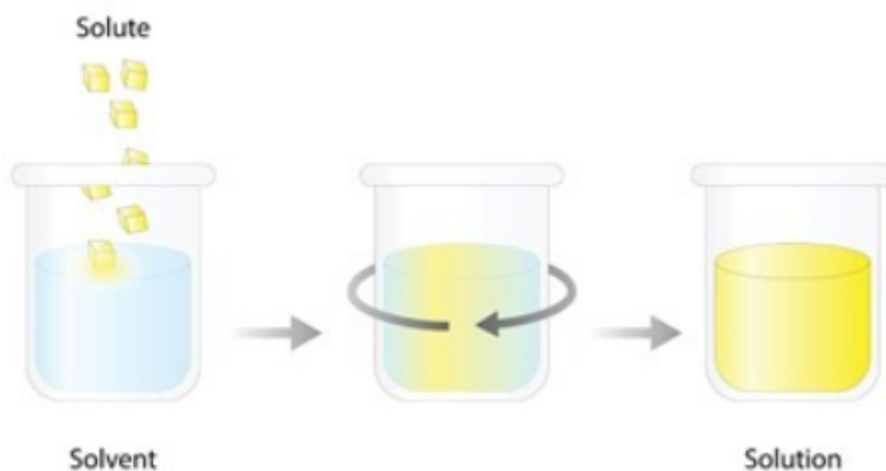
Solvents are primarily used for cleaning and dissolving substances, but they also play a key role in processes like paint thinning, adhesive removal, and chemical extractions. Whether working on a large industrial project or a simple DIY task, **solvents** make it easier to prep surfaces, clean up, and maintain product quality.

For example:

- Acetone is widely used for cleaning tools and removing adhesives.
- Ethanol is common in disinfectants and pharmaceuticals.
- Toluene helps dissolve paints and coatings.

While solvents are helpful in many applications, they can pose serious health risks if proper PPE isn't used. That's where MCR Safety steps in, providing gloves, goggles, and protective gear specifically designed to keep workers safe from solvent exposure.

Solute vs. Solvent



At the core of **understanding solvents** is knowing how they interact with solutes to form solutions.

- A solute is a substance that dissolves in a solvent. Solute can be solid, liquid, or gas. In a

homogeneous solution, the solute disperses evenly, while in a heterogeneous mixture, it remains unevenly distributed. The concentration of a solution depends on the ratio of solute to solvent.

- A solvent is the medium that dissolves the solute, most commonly a liquid but sometimes a gas or solid. Solvents break down solute particles, spreading them evenly throughout the solution. The amount of solute a solvent can dissolve depends on temperature and solubility.

Key Takeaways:

- Solute = the dissolved substance (e.g., salt in water)
- Solvent = the dissolving medium (e.g., water in saltwater)
- "Like dissolves like" = polar solvents dissolve polar solutes, and nonpolar solvents dissolve nonpolar solutes

Examples of Common Solvents and Solute

Understanding *real-world applications* makes these concepts more relatable:

Common Solvents:

- Ethanol ($\text{C}_2\text{H}_5\text{OH}$) – Found in disinfectants, perfumes, and fuel additives
- Methanol (CH_3OH) – Common in industrial processes and fuels
- Acetone ($\text{C}_3\text{H}_6\text{O}$) – A polar solvent widely used in nail polish removers and paint thinners
- Toluene (C_7H_8) – A nonpolar solvent used in paints and coatings
- Ethyl acetate ($\text{C}_4\text{H}_8\text{O}_2$) – Frequently used in adhesives and coatings

Common Solutes:

- Sodium chloride (NaCl) – Table salt, dissolves in water
- Sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) – Dissolves in water to create sweet solutions
- Carbon dioxide (CO_2) – Dissolves in liquid to create carbonation
- Oxygen (O_2) – Dissolves in water, essential for aquatic life

Solvent Hazards: Risks to Hands, Arms, and Eyes

Workers who handle solvents face significant health risks due to direct exposure and potential absorption, particularly to their hands, arms, and eyes. Solvents' chemical composition determines their hazard level, but many are highly toxic, corrosive, or irritating, posing both immediate and long-term dangers.

Primary Risks to Hands and Arms

- Skin Irritation & Chemical Burns – Solvents like acetone, toluene, and methyl ethyl ketone (MEK) can strip away natural skin oils, leading to dryness, cracking, and dermatitis. Stronger solvents may cause chemical burns.
- Absorption into the Skin – Some solvents, including benzene, methanol, and carbon disulfide, can penetrate gloves and skin, entering the bloodstream and causing systemic toxicity.
- Nerve Damage & Organ Toxicity – Hexane and methyl n-butyl ketone (MBK) can cause peripheral nerve damage, leading to permanent numbness or motor dysfunction. Repeated exposure may also harm the liver and kidneys.

Dangers to the Eyes

- Severe Eye Irritation – Solvents like xylene, methanol, and chloroform can cause intense burning, redness, and swelling.
- Chemical Burns & Corneal Damage – Exposure to strong acids, alkalis, or highly volatile solvents can result in permanent corneal damage or blindness.
- Vapor Exposure & Vision Impairment – Prolonged exposure to solvent vapors, especially toluene and methylene chloride, may lead to blurred vision, dizziness, and headaches.

Solvents pose serious risks to workers handling them, making the use of proper personal protective equipment (PPE)—such as chemical-resistant gloves, safety goggles, and protective sleeves—essential for workplace safety. Following safety protocols can also help prevent severe injuries.

Continue reading this blog in its entirety ***here*** to learn how MCR Safety can help keep your workforce safe with PPE for solvent safety.

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